

IGCSE Chemistry

Name: _____ Score: _____

1. Elements in the Periodic Table are arranged in order of increasing:

- ☐ proton (atomic) number
- ☐ mass only
- ☐ colour
- ☐ discovery date

2. How many electrons are in the outer shell of a Group I element?

3. How many electrons are in the outer shell of a Group VII element?

4. Which are properties of transition metals? (Select all that apply.)

- ☐ hard and strong
- ☐ high melting points
- ☐ form coloured compounds
- ☐ very low density

Tick every correct option.

5. Noble gases are unreactive because they have:

- ☐ a full outer shell of electrons
- ☐ one outer electron
- ☐ seven outer electrons
- ☐ no electrons

6. Elements in the same group have the same number of:

- ☐ outer-shell electrons
- ☐ protons
- ☐ neutrons
- ☐ shells

7. Going down Group I, reactivity:

- ☐ increases

- ☐ decreases
- ☐ stays the same
- ☐ becomes zero

8. Going down Group VII, reactivity:

- ☐ decreases
- ☐ increases
- ☐ stays the same
- ☐ doubles

9. The period number equals the number of electron shells.

- ☐ True ☐ False

10. Group I metals are soft enough to cut with a knife.

- ☐ True ☐ False

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1. **proton (atomic) number**

They are ordered by atomic number.

2. **1**

Group I elements have one outer electron.

3. **7**

Halogens have seven outer electrons.

4. **hard and strong; high melting points; form coloured compounds**

They are dense, not low-density.

5. **a full outer shell of electrons**

A full outer shell is very stable.

6. **outer-shell electrons**

Same outer electrons → similar properties.

7. **increases**

The outer electron is lost more easily down the group.

8. **decreases**

Halogens get less reactive down the group.

9. **True**

Each period adds a shell.

10. **True**

They are soft and have low densities.