

Question 25 (Answer: **C**)

- The inert or noble gases occupy Group VIII, the far right-hand column.
- They already have a full outer shell, so they have no tendency to react.
- So pick the element in the last column of the table.

Question 24 (Answer: **D**)

- Noble gases have a full outer shell, which is why they are so unreactive – statement 2 is right.
- They are monatomic, but they are *not* reactive, so statement 1 is wrong.
- Only one of the two statements is correct.

Question 21 (Answer: **C**)

- Both density and boiling point increase steadily down Group VIII.
- Krypton sits between argon and xenon in the table.
- So its density lies between 1.8 and 5.9 g/dm³ and its boiling point between –186 and –108 °C.

Question 23 (Answer: **D**)

- Noble gases are monatomic, not diatomic, so they do not form molecules.
- Helium has only two outer electrons, so "all have eight" is false.
- They are so unreactive that they do not form oxides at all.