Name:

IGCSE Chemistry: **w25 21**

22 Four metal oxides are listed.

1 Al_2O_3

2 CaO

3 CuO

4 ZnO

Which metal oxides are amphoteric?

A 1 and 4 **B** 1 only **C** 2, 3 and 4 **D** 2 and 3 only

IGCSE Chemistry: **w25 22**

21 Phosphorus is an element in Group V of the Periodic Table.

It burns in air to form an oxide that dissolves in water to form a solution with a pH of 1.

Which row describes this oxide of phosphorus?

	metal oxide	non-metal oxide	acidic oxide	basic oxide
A	✓	✗	✓	✗
B	✓	✗	✗	✓
C	✗	✓	✓	✗
D	✗	✓	✗	✓

key

✓ = yes

✗ = no

IGCSE Chemistry: **w25 23**

25 Which metal oxide reacts with aqueous sodium hydroxide to form a salt and water?

A calcium oxide

B copper(II) oxide

C sodium oxide

D aluminium oxide

IGCSE Chemistry: **s25 22**

20 Which three oxides are **all** acidic?

A CaO, NO_2, SO_2

B CaO, CO_2, Na_2O

C CO_2, NO_2, SO_2

D CO_2, Na_2O, SO_2

20 Which row identifies an acidic oxide, a basic oxide and an amphoteric oxide?

	acidic	basic	amphoteric
A	ZnO	SO_2	Al_2O_3
B	SO_2	Al_2O_3	Na_2O
C	ZnO	Al_2O_3	SO_2
D	SO_2	Na_2O	Al_2O_3

23 Element E is a metal in Group I of the Periodic Table and element G is a non-metal in Group VII. Both of these elements form oxides.

Which statement about their oxides is correct?

A Both oxides are acidic.

B Both oxides are basic.

C The oxide of E is acidic and the oxide of G is basic.

D The oxide of G is acidic and the oxide of E is basic.