

36 Ethene reacts with hydrogen in an addition reaction to form ethane.

Which catalyst is used for this reaction?

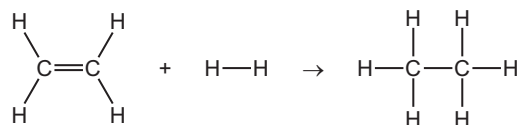
- A** iron
- B** nickel
- C** vanadium(V) oxide
- D** yeast

32 The alkenes are a homologous series.

Which statement about the members of the homologous series of the alkenes is correct?

- A** They have the same functional group.
- B** They have the same relative molecular mass.
- C** They have different chemical properties.
- D** They have different general formulae.

17 Ethene reacts with hydrogen to form ethane.



The bond energies are shown in the table.

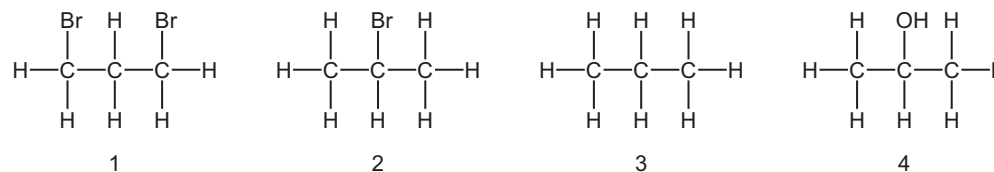
bond	bond energy in kJ/mol
C–C	+350
C–H	+410
H–H	+436
C=C	+614

What is the enthalpy change for the reaction?

- A** –290 kJ/mol
- B** –120 kJ/mol
- C** +120 kJ/mol
- D** +290 kJ/mol

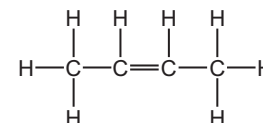
36 Propene undergoes addition reactions with bromine, hydrogen and steam.

Which structures represent a product of these reactions?

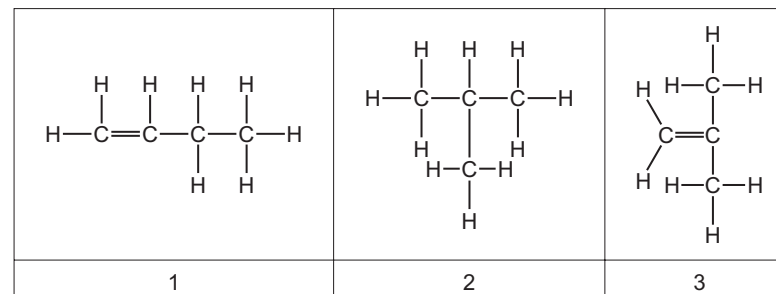


- A** 1 and 3 **B** 1 and 4 **C** 2 and 3 **D** 3 and 4

33 The displayed formula of but-2-ene is shown.



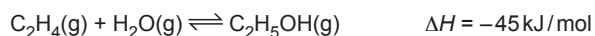
Which displayed formulae represent structural isomers of but-2-ene?



- A** 1, 2 and 3 **B** 1 and 2 only **C** 1 and 3 only **D** 2 and 3 only

6 Ethene, C_2H_4 , is the first member of a family of similar compounds which contains the alkene functional group.

- (a) State the term for a family of similar compounds which contain the same functional group.
..... [1]
- (b) Determine the difference in relative molecular mass between C_2H_4 and the next member in this family of similar compounds.
..... [1]
- (c) Write the symbol equation for the complete combustion of C_2H_4 .
..... [2]
- (d) C_2H_4 reacts with steam to form ethanol.



The process happens in a closed system and the reaction reaches an equilibrium.
The conditions for this process are 300 °C and 60 atm pressure. H_3PO_4 is used as a catalyst.

- (i) Complete Table 6.1 to show the effect, if any, on the concentration of $\text{C}_2\text{H}_4(\text{g})$ when the following changes to the conditions are applied.
- Only use the words **increases**, **decreases** or **no change**.

Table 6.1

change to condition	effect on the concentration of $\text{C}_2\text{H}_4(\text{g})$ at equilibrium
temperature is decreased	
some $\text{C}_2\text{H}_5\text{OH}(\text{g})$ is removed	
pressure is increased	
a more effective catalyst is used	

- (ii) Explain, in terms of collision theory, why the rate of the forward reaction increases if the temperature increases.

③

- (e) Compound **B** has the displayed formula shown in Fig. 6.1.

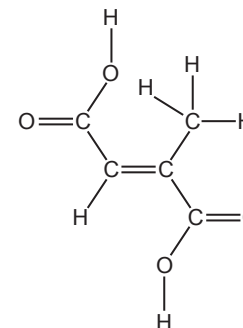


Fig. 6.1

- (i) Deduce the molecular formula of compound B.
..... [1]
- (ii) State why compound B is unsaturated.
..... [1]
- (iii) Draw the structure of one repeat unit of the polymer formed when compound B undergoes addition polymerisation.
- [2]
- (iv) Explain why 1 mol of compound B reacts with 2 mol of sodium hydroxide, NaOH.
.....
..... [1]
- (v) Calculate the volume, in cm³, of 0.250 mol/dm³ NaOH that reacts with 0.100 mol of compound B.

volume = cm³ [2]

The Periodic Table of Elements

Group																							
I	II	1 H hydrogen 1										III	IV	V	VI	VII	VIII						
		Key atomic number atomic symbol name relative atomic mass																					
3 Li lithium 7	4 Be beryllium 9																	5 B boron 11	6 C carbon 12	7 N nitrogen 14	8 O oxygen 16	9 F fluorine 19	10 Ne neon 20
11 Na sodium 23	12 Mg magnesium 24																	13 Al aluminium 27	14 Si silicon 28	15 P phosphorus 31	16 S sulfur 32	17 Cl chlorine 35.5	18 Ar argon 40
19 K potassium 39	20 Ca calcium 40	21 Sc scandium 45	22 Ti titanium 48	23 V vanadium 51	24 Cr chromium 52	25 Mn manganese 55	26 Fe iron 56	27 Co cobalt 59	28 Ni nickel 59	29 Cu copper 64	30 Zn zinc 65	31 Ga gallium 70	32 Ge germanium 73	33 As arsenic 75	34 Se selenium 79	35 Br bromine 80	36 Kr krypton 84						
37 Rb rubidium (87)	38 Sr strontium 88	39 Y yttrium 89	40 Zr zirconium 91	41 Nb niobium 93	42 Mo molybdenum 96	43 Tc technetium —	44 Ru ruthenium 101	45 Rh rhodium 103	46 Pd palladium 106	47 Ag silver 108	48 Cd cadmium 112	49 In indium 115	50 Sn tin 119	51 Sb antimony 122	52 Te tellurium 128	53 I iodine 127	54 Xe xenon 131						
55 Cs caesium 133	56 Ba barium 137	57–71 lanthanoids	72 Hf hafnium 178	73 Ta tantalum 181	74 W tungsten 184	75 Re rhenium 186	76 Os osmium 190	77 Ir iridium 192	78 Pt platinum 195	79 Au gold 197	80 Hg mercury 201	81 Tl thallium 204	82 Pb lead 207	83 Bi bismuth 209	84 Po polonium —	85 At astatine —	86 Rn radon —						
87 Fr francium —	88 Ra radium —	89–103 actinoids	104 Rf rutherfordium —	105 Db dubnium —	106 Sg seaborgium —	107 Bh bohrium —	108 Hs hassium —	109 Mt meitnerium —	110 Ds darmstadtium —	111 Rg roentgenium —	112 Cn copernicium —	113 Nh nihonium —	114 Fl flerovium —	115 Mc moscovium —	116 Lv livermorium —	117 Ts tennessine —	118 Og oganesson —						

anthanoids

ictinoids

the volume of one mole of any gas is 24 dm³ at room temperature and pressure (r.t.p.).

6 The structural formulae of two compounds, **A** and **B**, are shown.

A	B
CH ₂ =CHCH ₃	CH ₃ CH=CHCH ₃

A and **B** are members of the same homologous series.

(a) Give **two** reasons why the structural formulae of **A** and **B** show they are members of the same homologous series.

- 1
- 2 [2]

(b) Explain why **A** and **B** are both hydrocarbons.

..... [1]

(c) Write the symbol equation for the complete combustion of **A**.

..... [2]

(d) Deduce the empirical formula of **A**.

..... [1]

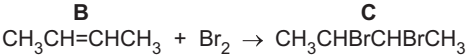
(e) Name compound **B**.

..... [1]

(f) A structural isomer of **B** is a member of the same homologous series.

Draw the displayed formula of this structural isomer of **B**.

(g) Compound **B** reacts with aqueous bromine at room temperature to form product **C**. The equation is shown.



(i) State why this is an addition reaction. [1]

(ii) Describe the colour change in aqueous bromine during this reaction. from to [1]

(iii) Name product **C**. [1]

(h) Under certain conditions, **one** mole of **B** reacts with oxygen to form **two** moles of carboxylic acid **D**.

Carboxylic acid **D** has **two** carbon atoms.

(i) Draw the displayed formula of carboxylic acid **D**. [1]

(ii) Name carboxylic acid **D**. [1]

(iii) Complete the symbol equation for this reaction. CH3CH=CHCH3 + → [1]

[Total: 14]

The Periodic Table of Elements

Group																	
I	II	Key										III	IV	V	VI	VII	VIII
		atomic number atomic symbol name relative atomic mass										1 H hydrogen 1					
3 Li lithium 7	4 Be beryllium 9											5 B boron 11	6 C carbon 12	7 N nitrogen 14	8 O oxygen 16	9 F fluorine 19	2 He helium 4
11 Na sodium 23	12 Mg magnesium 24											13 Al aluminium 27	14 Si silicon 28	15 P phosphorus 31	16 S sulfur 32	17 Cl chlorine 35.5	10 Ne neon 20
19 K potassium 39	20 Ca calcium 40	21 Sc scandium 45	22 Ti titanium 48	23 V vanadium 51	24 Cr chromium 52	25 Mn manganese 55	26 Fe iron 56	27 Co cobalt 59	28 Ni nickel 59	29 Cu copper 64	30 Zn zinc 65	31 Ga gallium 70	32 Ge germanium 73	33 As arsenic 75	34 Se selenium 79	35 Br bromine 80	36 Kr krypton 84
37 Rb rubidium 85	38 Sr strontium 88	39 Y yttrium 89	40 Zr zirconium 91	41 Nb niobium 93	42 Mo molybdenum 96	43 Tc technetium —	44 Ru ruthenium 101	45 Rh rhodium 103	46 Pd palladium 106	47 Ag silver 108	48 Cd cadmium 112	49 In indium 115	50 Sn tin 119	51 Sb antimony 122	52 Te tellurium 128	53 I iodine 127	54 Xe xenon 131
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87 Fr francium —	88 Ra radium —	89 Ac actinium —	90 Th thorium 232	91 Pa protactinium 231	92 U uranium 238	93 Np neptunium —	94 Pu plutonium —	95 Am americium —	96 Cm curium —	97 Bk berkelium —	98 Cf californium —	99 Es einsteinium —	100 Fm fermium —	101 Md mendelevium —	102 No nobelium —	103 Lr lawrencium —	104 Og oganesson —

anthanoids

actinoids

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