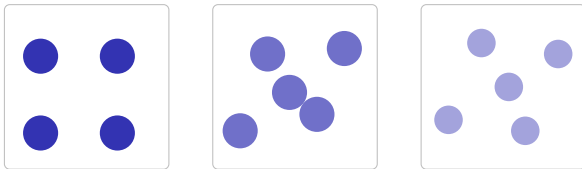


States of Matter

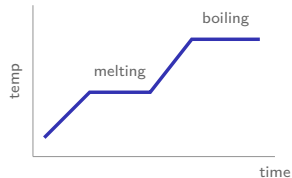
IGCSE Chemistry ■ Topic 1

IGCSE Chemistry



What we will cover

- 1 The three states & their particles
- 2 Changes of state
- 3 Heating & cooling curves
- 4 Gas pressure
- 5 Diffusion



heating and cooling curves



The three states of matter

- **solid** – regular; vibrate in place; fixed shape and volume
- **liquid** – close; can slide; fixed volume, takes the container
- **gas** – far apart, fast; fills the container; easy to compress

Kinetic particle theory : heating raises the particles' kinetic energy.

Properties you can observe

You do not need a microscope to tell the three states apart.

- A solid has a fixed shape and a fixed **volume** . It does not flow and you cannot **compress** it (squeeze it smaller).
- A liquid has a fixed volume but no fixed shape. It flows and takes the shape of its container. It is almost impossible to compress.
- A gas has no fixed shape and no fixed volume. It flows and spreads out to fill the whole container. A gas is easy to compress.

The three states, compared

the particles	atoms , molecules or ions – and they are always in motion
arrangement	solid: regular and touching. Liquid: close but disordered. Gas: far apart and random
density	how much mass is packed into a volume – high for solids and liquids, very low for gases
can it be compressed?	solid no ; liquid almost none ; gas yes, easily
does it flow?	solid no ; liquid and gas yes
fixed shape?	only the solid; a liquid keeps its volume, a gas fills its container

Every row follows from how far apart the particles are and how strongly they attract. Say that once and the whole table is derivable.

Changes of state

Melting and boiling take energy (to break attractions) at **constant** temperature. Freezing and condensing give that energy back.

Some solids **sublime** straight to gas (iodine, dry ice).

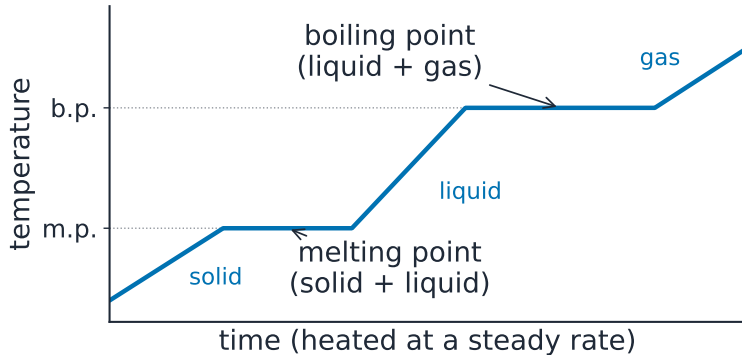
Heating and cooling curves

Two **flat** parts on a heating curve:

- at the **melting point** (solid + liquid)
- at the **boiling point** (liquid + gas)

The heat breaks **forces of attraction** , so the temperature holds steady.

Heating and cooling curves, in detail

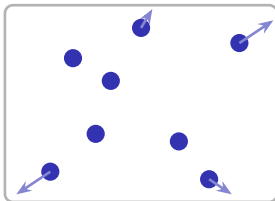


plateau: energy used as latent heat (bonds break) —
temperature stays constant

Temperature stays constant at the melting and boiling points while forces of attraction are broken



Gas pressure



particles hitting the walls
cause the pressure

heat → faster, harder, more
often → more pressure

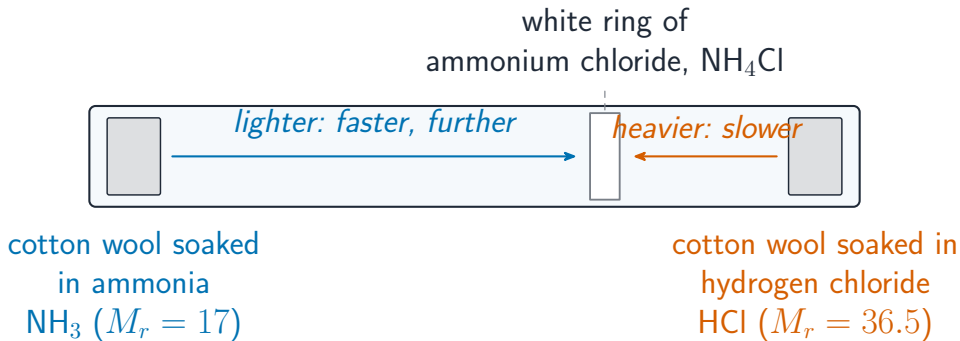
Squeeze a gas (smaller volume) and the particles hit the walls more often, so the **pressure** rises.

Diffusion

Diffusion : particles spread from high to low **concentration** .

A **lighter** gas diffuses faster. Light NH_3 ($M_r = 17$) beats heavy HCl ($M_r = 36.5$), so the white ring forms nearer the HCl end.

Diffusion, in detail



The **lighter** gas diffuses faster: NH_3 ($M_r = 17$) beats HCl ($M_r = 36.5$), so the white ring forms nearer the HCl end.

Worked example – which gas travels further?

Bromine ($M_r = 160$) and ammonia ($M_r = 17$) start at opposite ends of a long tube. Where do they meet?

- lighter particles move **faster** at the same temperature
- ammonia's **relative molecular mass** is far smaller, so it diffuses faster
- they meet **nearer the bromine end**, where the white ring of **ammonium chloride** forms

Rate of diffusion falls as M_r rises. The exam gives you two masses and expects you to say which travels further, **and why**.

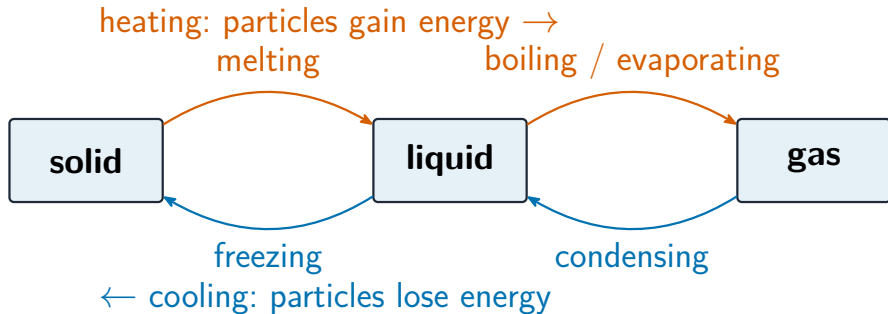
Particles diffuse from high to low concentration



start: **crowded** on one side end: spread out **evenly**
particles move randomly from **high** to **low** concentration - this is why you can smell food across a room

Particles diffuse from high to low concentration until they are spread out evenly

Changes of state, continued



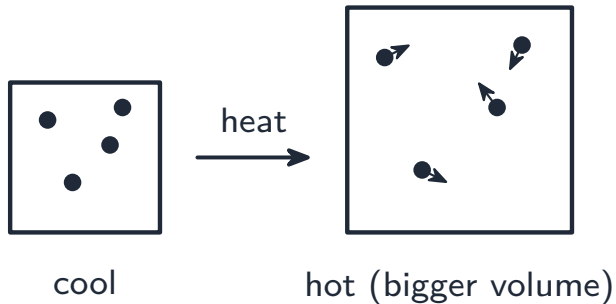
Heating: solid → liquid → gas; cooling reverses each change

Changes of state, continued (2)



Ice melting to water: a change of state as the solid warms.

Effect of temperature



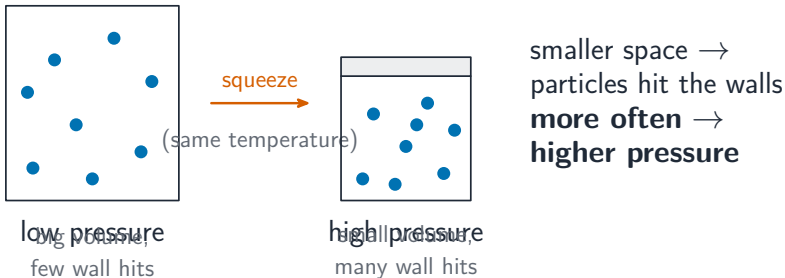
A heated gas: faster particles spread out and take up more space

Changes of state, continued (3)



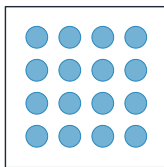
Warmed solid iodine turns straight into a purple gas (sublimation)

Effect of pressure



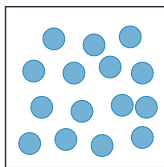
Squeezing a gas into a smaller volume makes the particles hit the walls more often, so the pressure rises

The particle picture



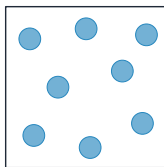
solid

regular, touching
vibrate in place



liquid

close, random
slide past
each other



gas

far apart,
random
move fast

Particles are packed and regular in a solid, close and random in a liquid, and far apart in a gas

Exam tips

- During **melting** and **boiling** the temperature stays constant (the flat parts of a heating curve), because the energy breaks the forces between particles instead of making them move faster.
- **Evaporation** happens only at the surface and at any temperature; **boiling** happens throughout the liquid at one fixed temperature.
- Explain gas behaviour with the particle theory: heating a gas at fixed volume raises the pressure because the particles hit the walls harder and more often – the particles themselves do not get bigger.
- **Diffusion** is faster for lighter particles (smaller M_r) and cannot happen in solids. In the ammonia/hydrogen chloride tube the lighter ammonia travels further, so the white ring forms nearer the HCl end.



Worked example – comparing diffusion

Ammonia ($M_r = 17$) and hydrogen chloride ($M_r = 36.5$) diffuse from opposite ends of a tube. Where does the white ring form?

- Diffusion is the random movement of particles from high to low concentration.
- At the same temperature all gas particles have the same average **kinetic energy**.
- NH_3 is **lighter**, so it moves **faster** than HCl .
- It travels further before they meet $\Rightarrow$ the ring forms **nearer the HCl end**.

Lighter particles diffuse faster **at the same temperature**. Say “lower $M_r \Rightarrow$ faster” – do not just say “it is a gas”.

Topic 1 – key takeaways

1

States

solid packed, liquid close, gas far apart

2

Changes

melt, freeze, boil, condense, sublime

3

Curves

flat at the melting & boiling points

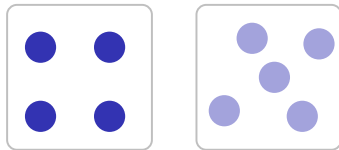
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Diffusion

lighter particles diffuse faster

Check yourself

- 1 In which state do particles vibrate in fixed places?
- 2 What is the name for solid turning straight to gas?
- 3 Why is a heating curve flat at the melting point?
- 4 Which diffuses faster: a light or a heavy gas?



Answers 1. a solid 2. sublimation 3. heat breaks the forces, not raising the temperature 4. the lighter gas