

5.2.1 Explain configurations, ligands and complex charge

Name: _____ Class: _____ Date: _____

Total: 16 marks

KEY WORDS evidence • conditions

Objective

Explain configurations, ligands and complex charge. All questions and marking guidance are original.

1 Key knowledge and worked examples

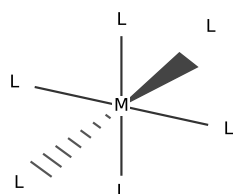
A transition element forms a stable ion with an incomplete d subshell. Remove 4s electrons before 3d when forming positive ions. A ligand donates an electron pair; coordination number counts directly attached donor atoms.

■ Worked method

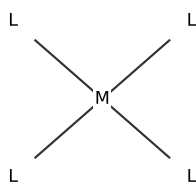
Fe is $[\text{Ar}]3\text{d}^64\text{s}^2$, so Fe^{2+} is $[\text{Ar}]3\text{d}^6$. In $[\text{Fe}(\text{CN})_6]^{3-}$, six CN^- contribute -6 ; iron must be $+3$. Three neutral bidentate en ligands give six donor atoms, so $[\text{Co}(\text{en})_3]^{3+}$ has coordination number six and cobalt(III). Colour depends on energy gaps between split d orbitals, not simply the total number of ligands.

Count donor atoms; then distinguish their spatial arrangement

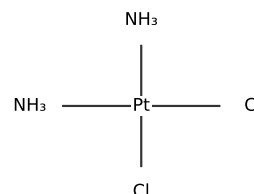
Octahedral: six donors



Square planar: four donors



Cisplatin: adjacent Cl



2 Practice

2.1 Give the ground-state configurations after [Ar] for Cr and Cu. [2]

2.2 Give the configurations of Fe^{2+} and Cu^{2+} after [Ar], explaining electron removal. [3]

2.3 Explain why zinc does not meet the transition-element definition but copper does. [2]

3 Exam-style questions

3.1 Determine metal oxidation state and coordination number in $[\text{Cr}(\text{H}_2\text{O})_4\text{Cl}_2]^+$ and $[\text{Co}(\text{en})_3]^{3+}$. Each en ligand donates through two nitrogen atoms. [4]

3.2 Explain why replacing water ligands with ammonia can change colour, and why a d^{10} ion often lacks d–d colour. [3]

3.3 Two complexes each have coordination number four. One is tetrahedral and one square planar. Explain why coordination number alone cannot distinguish them; identify the cis relationship in square-planar cisplatin. [2]
