

2.2.1 Intermolecular forces: drawings and evidence transfer

Name: _____ Class: _____ Date: _____

Total: 30 marks

KEY WORDS hydrogen bond

Objective

Draw and explain intermolecular attractions, then transfer the model to unfamiliar liquids and solvents. All questions and marking guidance are original.

1 Key knowledge and worked examples

A hydrogen bond joins a suitable donor H (covalently bonded to N, O or F) to a lone pair on N, O or F of another molecule. London forces act in all these liquids too. Polar molecules also have permanent dipole–dipole attractions; finding a dipole does not prove hydrogen bonding.

■ Worked comparison

Propanal and propan-1-ol both have three carbons. Propanal has a polar C=O bond and accepts hydrogen bonds from water, but has no N–H/O–H/F–H donor. Pure propan-1-ol has O–H donors and oxygen lone pairs, so its molecules form hydrogen bonds to one another. Compare particles and attractions; do not say boiling breaks the C=O or O–H bonds.

For solubility, compare attractions broken and formed. Water's O δ^- faces a cation; its H δ^+ faces an anion. Small alcohols form hydrogen bonds with water. A non-polar solvent can suit iodine because London attractions replace London attractions; polarity alone is not a complete solubility test.

2 Practice

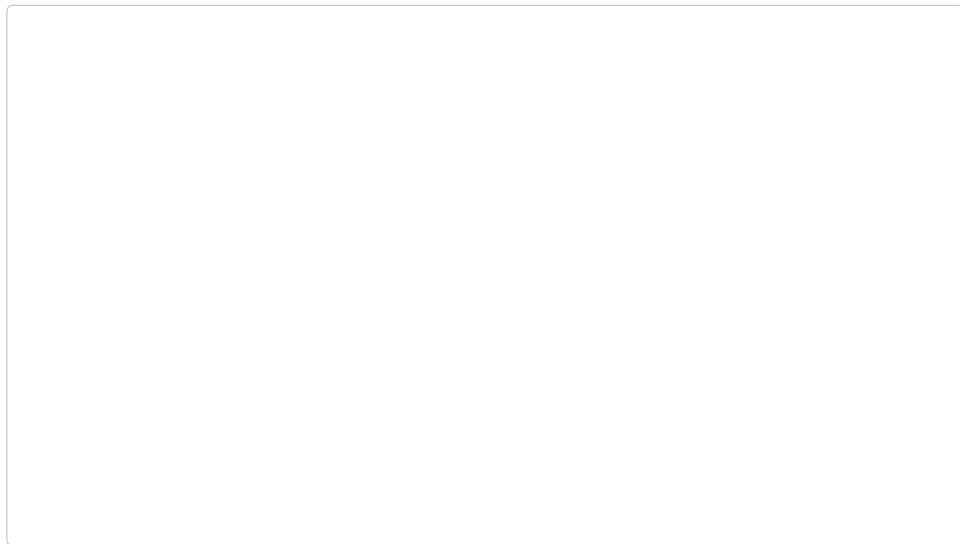
2.1 Compare pure propanal, CH₃CH₂CHO, and propane, CH₃CH₂CH₃. State the attractions each has, identify which has permanent dipole–dipole attractions, and explain why neither forms a hydrogen-bond network with its own molecules. [4]

2.2 Predict whether pure methanamide, HCONH₂, can hydrogen-bond to itself. Identify a donor and an acceptor. Compare pure propanone, CH₃COCH₃, and its ability to accept a hydrogen bond from water. [4]

2.3 Explain why boiling temperatures generally rise from methane to pentane. Then explain why a three-carbon alcohol usually boils above a similar-size alkane. State the type of bond or attraction chiefly overcome in each case. [4]

3 Exam-style questions

3.1 Draw a pair of molecules joined by one hydrogen bond for each liquid: H_2O , NH_3 and HF . Include every atom and covalent bond, all acceptor-atom lone pairs, the donor/acceptor partial charges and a distinct hydrogen-bond line. [6]



3.2 Rounded boiling temperatures in $^{\circ}\text{C}$ are HCl -85 , HBr -67 , HI -35 and HF $+20$. Explain the trend $\text{HCl} \rightarrow \text{HI}$ and the HF exception. Separately, compare water with the similar simple molecule H_2S (boiling -60°C): Water melts at 0°C , while H_2S melts at -85°C . Explain water's high melting and boiling temperatures, then why ice is less dense than liquid water near 0°C . [6]

3.3 Choose water or hexane as the more suitable solvent in each case: sodium chloride, ethanol, iodine. Explain the particle attractions. State how a water molecule orients toward Na^+ and Cl^- , and explain why a polar halogenoalkane may dissolve poorly in water. [6]
