

6 Finding enthalpy changes — Part 2 — Answers

Answers

Work through these after you have tried the questions yourself.

2.1 breaking a bond takes in energy; forming a bond gives out energy.

2.2 $\Delta H = 500 - 650 = -150$ kJ; it is exothermic (negative ΔH).

2.3 broken = $436 + 193 = 629$; formed = $2 \times 366 = 732$; $\Delta H \approx 629 - 732 = -103$ kJ.

2.4 its total ΔH equals the sum of the ΔH values of the individual steps, whatever route is taken.