

3.13 The Beer-Lambert Law

Name: _____ Class: _____ Date: _____

Total: 13 marks

KEY WORDS absorbance 吸光度 • solution 溶液

Objective

Build the skills to answer exam questions on the **Beer-Lambert law** — linking absorbance to concentration.

You must be able to:

- use $A = \epsilon bc$ (**absorbance** 吸光度)
- interpret a calibration plot of A vs concentration
- find an unknown concentration from absorbance

1 Worked examples

Study these first. Each one shows the method for a question type used later — follow the steps and you can do the Practice and Exam-style questions yourself.

■ The Beer-Lambert law

$$A = \epsilon bc,$$

where A is absorbance, ϵ the molar absorptivity, b the path length, and c the concentration. Absorbance is **proportional to concentration** (fixed ϵ and b).

■ A worked value

If $\epsilon = 200 \text{ L mol}^{-1}\text{cm}^{-1}$, $b = 1.0 \text{ cm}$, $c = 0.0020 \text{ M}$:

$$A = (200)(1.0)(0.0020) = 0.40.$$

■ The calibration line

Plot A (y) against concentration (x): a straight line through the origin with slope ϵb . Read an unknown concentration by finding its absorbance on the line.

■ Finding an unknown

If a sample has $A = 0.30$ and the calibration line is $A = 150c$, then $c = \frac{0.30}{150} = 0.0020 \text{ M}$.

2 Practice

Now apply the methods above.

2.1 Write the Beer-Lambert law and name each symbol. [2]

2.2 Find A for $\varepsilon = 500$, $b = 1.0$ cm, $c = 0.0010$ M. [2]

2.3 Absorbance is proportional to which quantity? [1]

3 Exam-style questions

3.1 In $A = \varepsilon bc$, doubling the concentration c (with ε , b fixed) [1]

- **A** halves A
- **B** doubles A
- **C** leaves A unchanged
- **D** squares A

3.2 A calibration line for a dye is $A = 250c$ (with c in mol/L).

(a) Find the absorbance of a 0.0016 M solution. [2]

(b) A sample gives $A = 0.50$. Find its concentration. [2]

3.3 A solution's absorbance is measured as 0.80 with $\varepsilon = 400$ L mol⁻¹cm⁻¹ and $b =$

1.0 cm. Find the concentration.

[3]

Solutions

2.1 $A = \varepsilon bc$: A absorbance, ε molar absorptivity, b path length, c concentration.

2.2 $A = (500)(1.0)(0.0010) = 0.50$.

2.3 Concentration.

3.1 B — A is proportional to c , so doubling c doubles A .

3.2 (a) $A = 250(0.0016) = 0.40$. (b) $c = \frac{0.50}{250} = 0.0020 \text{ M}$.

3.3 $c = \frac{A}{\varepsilon b} = \frac{0.80}{(400)(1.0)} = 0.0020 \text{ M}$.