

34 Basicity and phenylamine — Part 2

— Answers

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Work through these after trying the questions yourself.

34.6 The lone pair on the nitrogen.

34.7 An H^+ (a proton).

34.8 An amine with $-\text{NH}_2$ joined to a benzene ring.

34.9 Nitrate benzene to nitrobenzene.

34.10 Tin (Sn) and concentrated HCl.

34.11 In phenylamine the nitrogen's lone pair is partly delocalised into the benzene ring, so it is *less available* to accept an H^+ . In ethylamine the alkyl group instead pushes electron density onto the nitrogen, making its lone pair *more available* — so ethylamine is the stronger base.