31 Why halogenoarenes are unreactive – Part 2 — Answers

Answers

Work through these after trying the questions yourself.

31.6 Nucleophiles (by nucleophilic substitution).
31.7 It overlaps the delocalised ring.
31.8 It becomes shorter and stronger (harder to break).
31.9 The electron-rich ring repels the nucleophile.
31.10 No.

31.11 (1) The chlorine's lone pair overlaps the delocalised ring, giving the C–Cl bond partial double-bond character, so it is shorter and stronger and hard to break. (2) The electron-rich ring repels the approaching nucleophile, so it cannot attack the carbon.