

## 26 Rate equations and orders — Part 1

### — Answers

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#### Answers

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Work through these after trying the questions yourself.

**26.1** How the rate depends on the reactant concentrations.

**26.2** The order of reaction with respect to A.

**26.3** The sum of the individual orders ( $m + n$ ).

**26.4** By experiment.

**26.5** Zero order.

**26.6** The order with respect to B is 2. Since  $\text{rate} \propto [\text{B}]^2$ , doubling [B] multiplies the rate by  $2^2 = 4$  — the rate becomes four times larger.