

24 Electrolysis – Part 1 — Answers

Answers

Work through these after trying the questions yourself.

24.1 Using electricity to break down a molten or aqueous electrolyte.

24.2 Positive ions gain electrons (reduction).

24.3 Negative ions lose electrons (oxidation).

24.4 The metal.

24.5 Because it also contains water, which can be discharged instead.

24.6 Cathode: lead metal, $\text{Pb}^{2+} + 2\text{e}^- \rightarrow \text{Pb}$. Anode: bromine, $2\text{Br}^- \rightarrow \text{Br}_2 + 2\text{e}^-$.